

Electron Configuration Practice
Chemistry

Name: _____

Date: _____

A. In the space below, write the full (unabbreviated) electron configurations.

1. In: _____

2. Te: _____

3. Sr: _____

4. Kr: _____

5. Ba: _____

In the space below, write the Noble Gas (abbreviated) electron configurations.

6. Ag: _____

7. Ti: _____

8. C: _____

9. Cl: _____

10. Mo: _____

11. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$: _____

12. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$: _____

13. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$: _____

14. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$: _____

15. $[\text{Kr}] 5s^2 4d^{10} 5p^3$: _____

16. $[\text{Ar}] 4s^1$: _____

17. $[\text{Xe}] 6s^2$: _____

18. $[\text{Ne}] 3s^2 3p^1$: _____

19. $[\text{He}] 2s^2 2p^6$: _____

20. $[\text{Kr}] 5s^2 4d^8$: _____

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B. In the space below, write the full (unabbreviated) electron configurations of the following elements:

- 1) sodium _____
- 2) iron _____
- 3) bromine _____
- 4) barium _____
- 5) caesium _____

C. In the space below, write the Noble Gas (abbreviated) electron configurations of the following elements:

- 6) cobalt _____
- 7) silver _____
- 8) tellurium _____
- 9) rubidium _____
- 10) oxygen _____
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ _____
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ _____
- 13) $[\text{Kr}] 5s^2 4d^{10} 5p^3$ _____
- 14) $[\text{Xe}] 6s^2 4f^{14} 5d^6$ _____
- 15) $[\text{Kr}] 5s^2$ _____